



Intro to Vacuum Technology

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Gas vs. Liquid

Gas: $1 \text{ atm} = 760 \text{ torr (mm Hg)}$
 $= 760,000 \text{ "microns" (}\mu\text{m Hg)}$
 $= 101 \text{ kPascals (kPa)}$
 $= 1.01 \text{ Bar (B)}$
 $= 2.4 \times 10^{22} \text{ particles/L}$
 $= 2.4 \times 10^{-2} \text{ particles/nm}^3$
 $\equiv 50 \text{ nm}^3/\text{particle}$ $\text{mfp} \sim 65 \text{ nm}$

Liquid (1000nm) $\rho \approx 1 \Rightarrow 10 \text{ M}$
 $= 6 \times 10^{24} \text{ particles/L}$
 $\equiv 0.6 \text{ nm}^3/\text{particle}$

300X

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Vacuum

Vacuum

$15 \text{ torr} = 2 \times 10^{-4} \text{ atm} = \text{aspirator pump}$

$10^{-6} \text{ atm} \sim 1 \mu\text{mHg} \Rightarrow \text{"roughing" pump}$

$10^{-9} \text{ atm} \Rightarrow \text{diffusion pump "High Vac"}$

$10^{-12} \text{ atm} \Rightarrow \text{Ti getter / Light Cryopump "Ultra-high Vac"}$

$\hookrightarrow 5 \times 10^{15} \text{ nm}^3 / \text{particle}$
 $\sim 0.1 \text{ mm cube / particle}$

$m_p \leftarrow 6m$
 "molecular flow"

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Vacuum

TABLE 3.1
Common Vacuum Units

1 torr (T) = 1 mm Hg

1 micron, or micrometer (μm) = 10^{-3} torr

1 torr = $1/760$ atm

1 pascal (Pa) =
1 newton/ $\text{m}^2 = 7.5 \times 10^{-3}$ torr

1 pascal = $1/101,325$ atm

1 millibar (mbar) =
 10^3 dynes/ $\text{cm}^2 = 0.75$ torr

1 millibar = $1/1,013$ atm

TABLE 3.2
Qualitative Vacuum Ranges

Laminar Flow Range

Low or rough vacuum	1 atm–100 Pa
Medium or fine vacuum	100–0.1 Pa

Molecular Flow Range

High vacuum	0.1– 10^{-4} Pa
Very high vacuum	10^{-4} – 10^{-7} Pa
Ultra-high vacuum	$\leq 10^{-7}$ Pa

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Rotary Pump

Figure 3.1 Rotary pumps. (A) Schematic of a rotary-vane pump. (B) Schematic illustrating the concept of staging. The two stages are usually mounted coaxially to decrease the number of moving parts. Rotary pumps are used more frequently than any other type of vacuum pump in the EM lab.

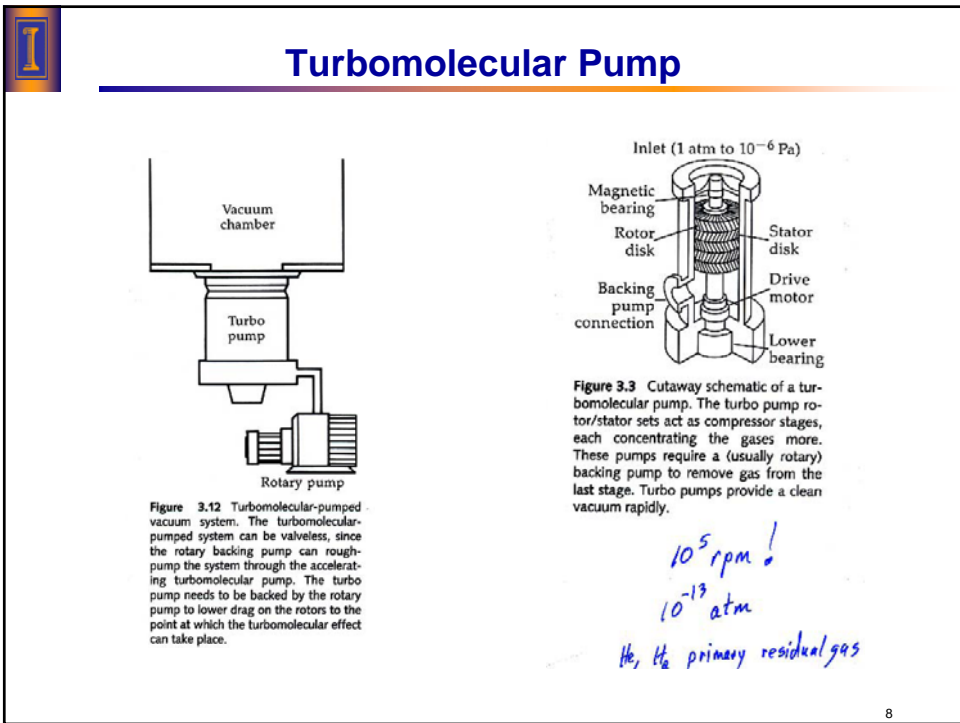
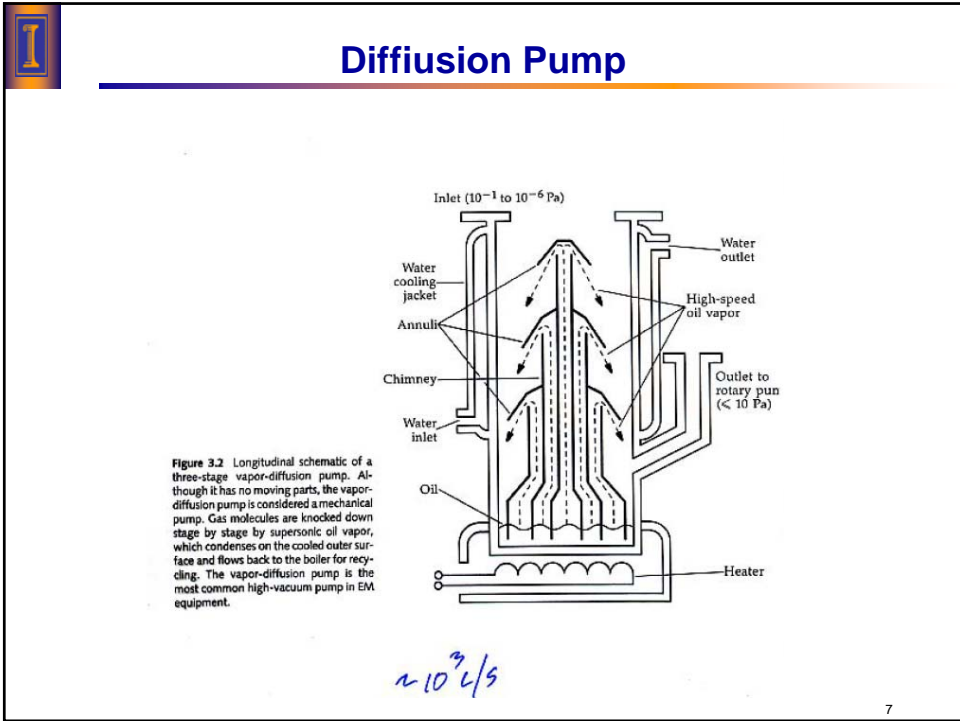
100-1000 L/min

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Diffusion Pump

Figure 3.11 Diffusion-pumped vacuum system. This vacuum system is the one found on most electron microscopes. The rotary pump first rough-pumps the chamber or column; then the final vacuum is attained by opening the main valve to the diffusion pump. The rotary pump then backs the diffusion pump by removing gas compressed in the space of its bottom stage.

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Turbomolecular Pump

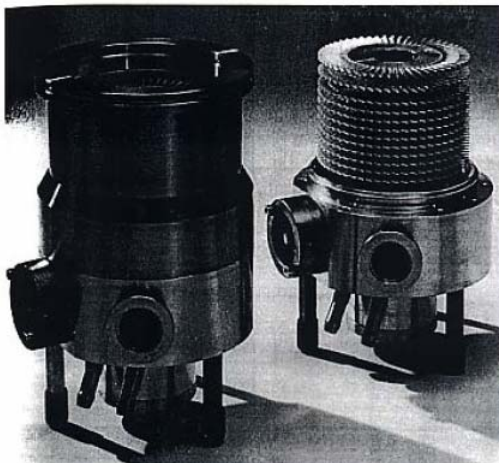


Figure 8.3. A turbopump (with and without its casing) which is nothing more than a small turbine that rotates at high speed. Like a jet turbine it pulls air in at the front end and forces it out of the back. The blades are designed like airfoils to enhance the flow of gas through the system.

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Ti Getter/Sublimation Pump

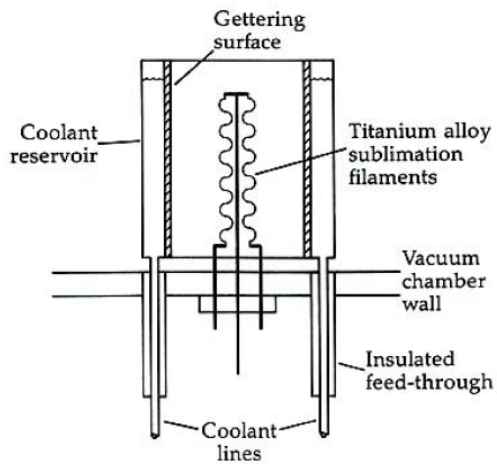


Figure 3.4 Longitudinal schematic of a titanium-sublimation pump. Titanium-sublimation pumps are found on ultrahigh-vacuum systems. Gas molecules stick to the walls of the pump body freshly coated with titanium. These surfaces are much more efficient at lower temperatures.

$\sim 10^4$ /s

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Ti Getter/Sublimation Pump

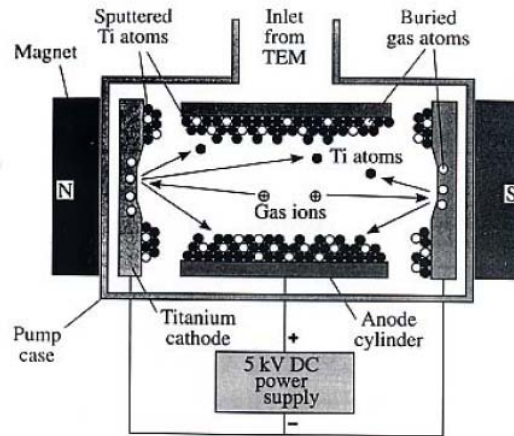


Figure 8.4. Schematic diagram showing how ion pumps trap ionized gas atoms by layers of Ti atoms at electrodes. Once trapped, the ions cannot escape until the pump is turned off.

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Cryo - Pump

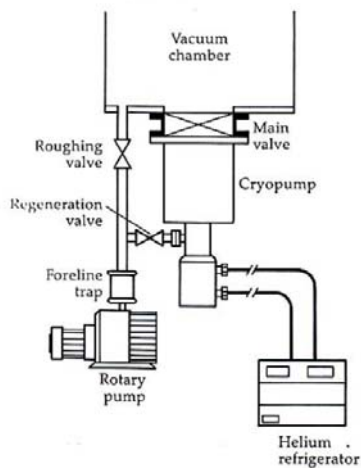


Figure 3.13 Cryo-pumped vacuum system. The cryopumped system requires a roughing pump, much like the diffusion-pumped system. Once the bulk of the gas has been rough-pumped out of the system and the plate valve to the cryoadsorption pump opened, the backing pump is no longer needed, since the gas is entrained in the cryopump head. As the adsorptive surfaces become saturated, the pump needs to be regenerated by heating the pump head and pumping out the liberated gas with the rotary pump.

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Cryo - Pump

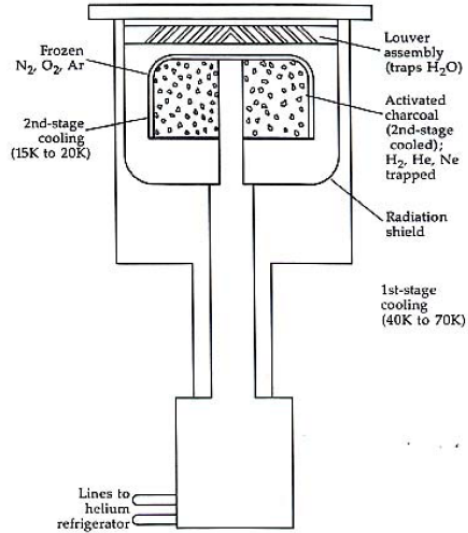


Figure 3.6 Longitudinal schematic of a cryoadsorption pump head. The two-stage cryoadsorption pump uses two cooling ranges either to freeze or to adsorb gas. The first stage traps water vapor and shields the inner stages. The second stage traps all gases either by freezing them or by adsorbing them.

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Vacuum Gauges

TABLE 3.4
Vacuum Gauge Ranges

Type of Gauge	PRESSURE RANGE (Pa)							
	10^4	10^2	10	10^{-2}	10^{-4}	10^{-6}	10^{-8}	10^{-10}
Thermocouple	-----							
Pirani	-----							
Cold-cathode				-----				
Hot-cathode					-----			

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Thermocouple Gauge

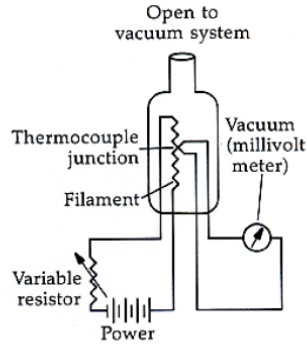


Figure 3.8 Thermocouple gauge system. Schematic of the gauge tube and the measuring circuit. The thermocouple gauge simply measures the temperature on a heated wire with a thermocouple, which is attached to it. As the vacuum increases, the wire gets hotter because fewer gas molecules are colliding with it and removing heat.

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Cold Cathode Gauge

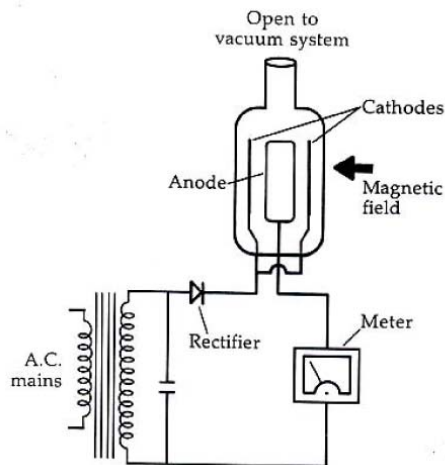


Figure 3.9 Cold-cathode gauge system. Schematic of the gauge tube and the measuring circuit. The cold-cathode gauge works by a magnetic field that lengthens the time of flight of electrons emitted from a cold source, improving the chance that they will hit a gas molecule and ionize it. Gas molecules ionized by these collisions create a current proportional to the vacuum.

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Hot Cathode Gauge

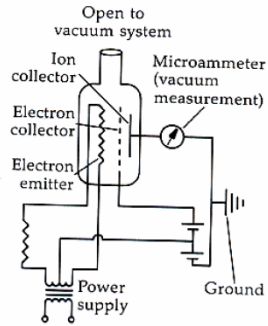


Figure 3.10 Hot-cathode gauge system. Schematic of the gauge tube and the measuring circuit. At higher vacuums, the cold discharge of electrons is insufficient. By heating the electron emitter, the electron flux improves, providing a higher probability that gas molecules will be ionized and eliminating the need for a magnetic field. Hot-cathode gauges can measure higher vacuums but are more sensitive to leaks than cold-cathode gauges are.